

MULLIKEN'S THEORY IN CHARGE-TRANSFER

COMPLEXATION

Dr.T.Sujatha¹, Reshma Jabeen²

¹Professor of Chemistry, HOD-Science & Humanities

Sri Sarathi Institute of Engineering & Technology, Nuzvid, Krishna Dt. A.P

³Academic Learner, Department of Sciences, The Hindu College

Machilipatnam, Krishna Dt.A.P

ABSTRACT

The nature of the attraction in a charge-transfer complex is not a stable chemical bond and is much weaker than covalent forces, rather it is better characterized as a weak electron resonance. As a result, the excitation energy of this resonance occurs very frequently in the visible region of the electro-magnetic spectrum. This produces the usually intense colors characteristic for these complexes. The optical absorption bands are often referred to as charge-transfer bands, or CT bands. Therefore optical spectroscopy is a powerful technique to characterize charge-transfer bands. The CT-complexes extend from hydrogen bonding to proton transfer reactions including solvent polarity scales based on them. These complexes are becoming increasingly important in all fields of human endeavour from physics and chemistry and biology to medicines. Currently there is an opinion among some scientists that these complexes are involved somewhere or the other on the reaction profile of all chemical reactions. They are, therefore, a special case of molecular association and have been most widely investigated. Charge transfer complexes are of immense importance in material science and they play a significant role in drug design and bioelectrochemical processes.

Keywords: acceptor, biochemical, electron, energy, molecules, solvent, transfer

Charge-transfer complexation is important phenomenon in biochemical and bioelectrochemical energy transfer processes [25]. Charge-transfer phenomenon was introduced first by Mulliken [26] and widely discussed by Foster to define a new type of adducts. Mulliken defines a molecular complex between two molecules as an association somewhat stronger than ordinary. Van der Waal's association of definite stoichiometry (1:1 for most cases). The partners are very often already have closed-shell (Saturated Valence) electronic structure. In loose complexes the identities of the original molecules are to a large extent preserved.

The tendency to form complexes occurs when one partner is an electron acceptor and the other is an electron donor. We abbreviate the term donor-acceptor complex to include all such associations and use D for an electron donor and A for electron acceptor. Most studies of complexes thus far have been made in solution, in solvent that are as inert as possible. It may, therefore, be assumed that the London dispersion interactions which are important between D and A in the vapour state, are very approximately cancelled by losses of solute-solvent dispersion force attractions when complex is formed from free donor and acceptor in solution. Roughly, one

donor- solvent plus one acceptor- solvent contact is replaced by one donor-acceptor and one solvent-solvent contact. The theory of donor-acceptor complexes and their spectra as presented by Mulliken is a vapour-state theory, except for the omission of the London dispersion attraction terms. This theory after small correction for solvation energies, is essentially valid for solutions in inert solvents. The few studies that have been made in vapour-state complexes are in agreement with this theory but they show some puzzling features.

The new bond formed according to Benesi and Hildebrand is in the ultra-violet region for a solution of benzene and iodine dissolved in n-heptane. Similar bonds also occur in the visible region for many other complexes.

This demonstrates that a solution of tetracyanoethylene (TCNE) in methylene dichloride may be added to a series of aromatic hydrocarbons dissolved in methylene dichloride, benzene gives a yellow solution, xylene an orange, durene a deep red and hexamethyl benzene a deep purple and a green colour was obtained on interaction of colourless solution of the TCNE acceptor and purple solution of porphyrins in dichloromethane or carbon tetrachloride describing a charge transfer complex (CTC) formation [27] and also when tetracyanoethylene solution was added to the solution of 2,6-diaminopyridine (2,6-DAPY), strong change in colour was observed and associated with the appearance of new absorption band [28] in region where neither donor nor acceptor have any absorption. The π -electron molecules, ethylene and benzene can act either as weak donors or weak acceptors. Other things being equal, donor ability increases with decreasing ionization potential (I_D); the acceptor ability with increasing electron affinity (E_A).

Among aromatic hydrocarbons I_D decreases and E_A increases with increasing size; graphite with $I_D = E_A$ is the extreme example and is in fact both a good acceptor and a good donor. Starting with any unsaturated or aromatic hydrocarbon, either its donor or its acceptor capability can be strengthened by the introduction of suitable substituent groups. The weak donor properties of benzene are fortified by adding more and more electron-releasing methyl groups. The two kinds of molecular

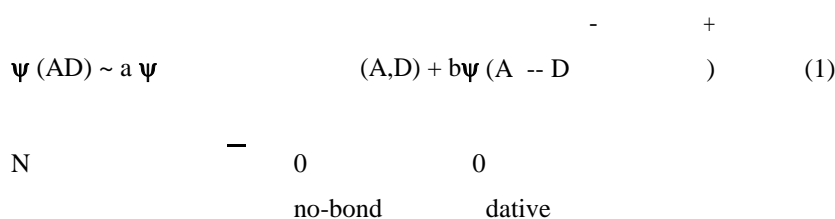
Complexes discussed above provide examples of n.v. (strong) and π - π^* (weak) complexes. The common types of donors and acceptors are listed in Table 1.

G.N. Lewis [29] explained co-ordination compounds or dative compounds

(e.g., $R_3N:BCl_3$ which can also be considered essentially as a stable molecular

complex) in terms of a structure with sharing of the electron lone pair of the nitrogen atom as well as the B (Boran) atom are surrounded by a complete octet of outer-shell electrons. This sharing can be expressed in quantum language by an approximate wave function ψ that is a combination of two resonance structure D (R_3N) and A

(BCl_3):



The dative structure corresponds to an ionic plus a covalent bond and has been called sometimes as semipolar double bond. The interpretation of the N-B dative bond in the complex, given by equation (1), is analogous to the approximate ionic covalent resonance interpretation of the chemical bond in HCl:

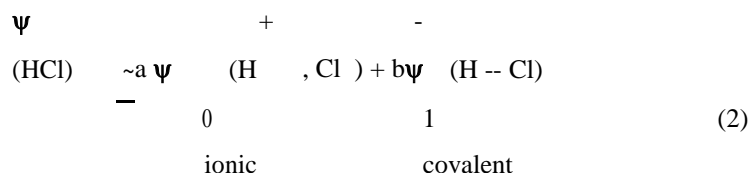


Table 1

Donor Type	Example	Dative Electron ^a from	Acceptor Type	Example	Dative Electron ^a goes to
n	:NR ₃	Non-bonding lone pair	V	BCl ₃	Vacant orbital
bπ	Benzene	Bonding π-orbital	aπ	TCNE	Antibonding π-orbital
			aσ	I ₂ , HQ ^b	Antibonding σ-orbital

a. —Dative electron|| refers to the electron transferred from donor to acceptor.

b. Molecules such as phenol, water, and other molecules that give hydrogen bonding.

In both examples b>>a.

The inclusion of the no-bond structure in equation (1) is given more importance than that of the ionic structure in equation (2).

Complexes are classified as strong or weak depending on whether the energy of formation and the equilibrium constant (K) are large or small. Relatively strong donors and/or acceptors form relatively stable complexes.

K



Equation (1) shows that the complex is stabilized by resonance between ψ_0 and ψ_1 . The forces involved being called charge-transfer (CT) forces. However, classical electrostatic forces (including induction forces) also contribute to the stability of the complexes and may even be of predominant importance for the stability of the most hydrogen bonded complexes and of the weaker of the complexes of the $b.\pi - a.\sigma$ and the benzene-iodine ($b.\pi - a.\sigma$) type [30]

In terms of the resonance structure description of (1) the structure of the ground state of any 1:1 complex is

$$\psi_N = a\psi_0(A,D) + b\psi_1(A \text{ -- } D)$$

This function is normalized as follows:

$$\int \psi_N \psi_N d\tau = \int a^2 \psi_0^2 d\tau + \int b^2 \psi_1^2 d\tau + 2ab \int \psi_0 \psi_1 d\tau \quad (3)$$

When ψ_N, ψ_0 and ψ_1 are normalized, then:

$$\int \psi_N^2 d\tau = 1, \quad \int \psi_0^2 d\tau = 1 \quad \text{and} \quad \int \psi_1^2 d\tau = 1$$

$$\int \psi_N^2 d\tau = 1, \quad \int \psi_0^2 d\tau = 1 \quad \text{and} \quad \int \psi_1^2 d\tau = 1$$

Substituting these values in equation (3) we obtain

$$a^2 + 2ab \int \psi_0 \psi_1 d\tau + b^2 = 1 \quad (4)$$

$\int \psi_0 \psi_1 d\tau$ is the so called the overlap integral and is represented by S_{01} with the integration carried over all space. If the complex is loose S_{01} is small and

$$a^2 + b^2 - 2ab S_{01} \sim 1 \quad (5)$$

Here, b approximately measures the weight of the dative structure or the fraction of an electron transferred from the donor to the acceptor in the ground state.

The term $2abS_{01}$ can easily be as large as b or even larger. Half of this term can reasonably be assigned to the donor and half to the acceptor so that the fractions F_0 and F_1 in the no-band and dative structure are:

2
2

F0

= a

+ abSo1, F1

= b

+abSo1

(6) In loose complexes between

2

2

closed -shell donors and acceptors $b \ll a$

2

For benzene. I_2 , I is approximately 0.06 or less; for pyridine. I_2 , b is approximately 0.2; for trimethylamine. I_2 , b may be 0.4.

If the ground state structure of the complex (weak or strong) is given by ψ_N then according to quantum theory principles, there must be an excited state ψ_E where ψ_E refers to the CT state. ψ_E is given by

$$\psi_E(AD) = a^* \psi_1(A - D) + b^* \psi_0(A, D) \quad (7)$$

The coefficient a^* and b^* are determined by the quantum theory requirement that the excited state wave function be orthogonal to the ground state function i.e.

$\int \psi_N \psi_E d\tau = 0$. The excited state function ψ_E is normalized as follows:

$$\int \psi_E^2 d\tau = a^{*2} + b^{*2} - 2a^*b^*S_{01} = 1 \quad (8)$$

This makes $a^* \sim 1$ and $b^* \sim 0$. If S_{01} were zero, $a^* = 1$ and $b^* = 0$ would be true

exactly.

For loose molecular complexes the ground state is mostly no-bond state that is,

$a \gg b$

. According to the orthogonality requirements the excited state is mostly

electron from ψ_N to ψ_E essentially dative i.e. $a^* \gg b^*$. The excitation of an

amounts to the transfer of an electron from D to A. The theory further shows that spectroscopic absorption from ψ_N to ψ_E should occur with generally high intensity.

Complexes have been studied mostly in solution but some studies have been made in solids [31-43] and in the case of (non-bonded vacant orbital) n.v. compounds and also recently of a few complexes of weaker types in the vapour state [44]. Complexes in solids even with 1:1 stoichiometry do not always occur in the form of pairwise units. Studies in the vapour state are difficult because K is small and interference of overlapping spectra of the

uncomplexed components is often severe. These difficulties are also found for solution studies but they are less troublesome because K is larger.

The complete absorption spectrum of a complex consists of absorption due to the following:

- (1) Locally excited states (states of A or of D, more or less but usually not greatly modified in the complex).
- (2) CT States [(Ψ_E) , as in eqn. (7) and other CT states involving excited dative structures, for example $\Psi_{(D^+ \leftarrow A^-)}$, $\Psi_{(D^+ \leftarrow A^*)}$].

Fig. 1.1 shows the change that occurs in the spectrum of iodine when it is dissolved in n-heptane and then when ethanol is added (Ethanol is transparent upto 220 nm). Iodine is well known for its electron-accepting properties, which may be deduced from molecular orbital consideration. It has been used in the past as a model acceptor to investigate the electron-donating properties of organic molecules [45] and during the past few decades the charge-transfer complexation of iodine with a wide variety of drugs molecule has been the subject of extensive research [46-49].

The maxima of the $C_2H_5OH.I_2$ CT band is marked in the Fig. (1.1), and the position of the shifted visible absorption band of I_2 in the complex (a transition to a locally excited state) is also indicated. The contact CT band appears as a long wavelength shoulder on the ultraviolet iodine band when I_2 is dissolved in heptane. It is felt by some authors that the importance of CT forces in stabilizing the ground state of such complexes has been exaggerated.

Multiple Charge-transfer Spectra

We can see from Table 2 that the intensity of CT spectra of the iodine-benzene complex is very large and that both the K value and the value of $-\Delta H$ increases

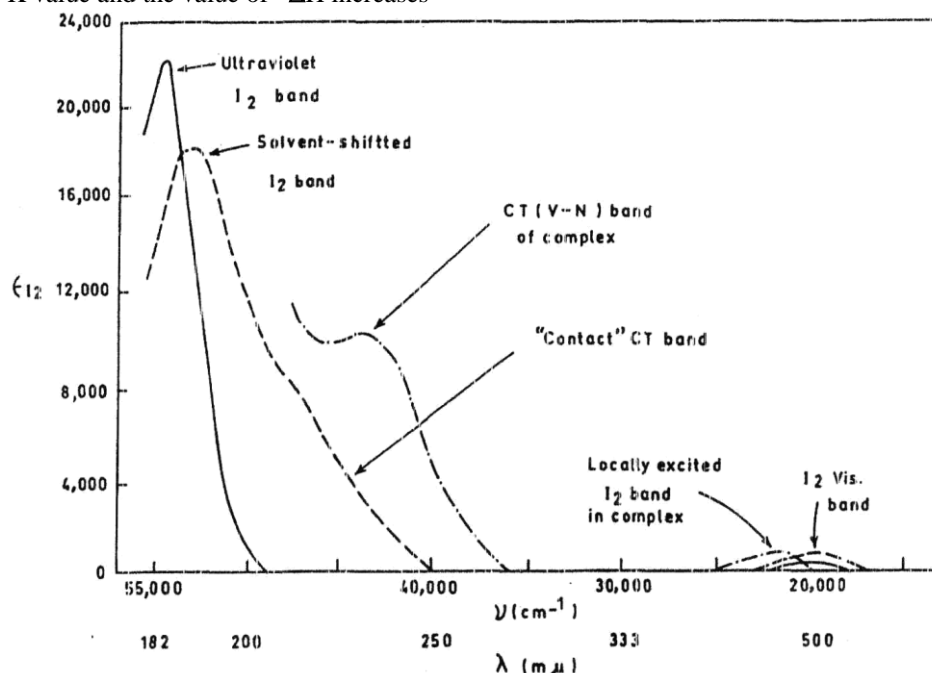


Fig. 1.1. The apparent molar absorptivity of I_2 vapour and of I_2 and EtOH. I_2 in n-heptane, Here (A) is for I_2 vapour; (B) for I_2 in n-heptane, and (C) for I_2 in n-heptane with 3.4 M ethyl alcohol.

Table 2

Spectroscopic and thermodynamic data of ionic complexes with some aromatic hydrocarbons

Donor	K _C (Temp.) (Solve)	ε _{CT} (λ _{max} CT) (Solve)	ΔH (Solv) K cal mole ⁻¹
Benzene	0.15 (25 ⁰ C) ^a CCl ₄	16.400 (292 mμ) (CCl ₄) ^a	-1.3 (Hexane) ^b (CCl ₄) ^c
Naphthalene	0.25 (25 ⁰ C) ^a CCl ₄	7.150 (360 mμ) (hexane) ^b	-1.8 (Hexane) ^b
Phenanthrene	0.45 (23 ⁰ C) ^d CCl ₄	7,100 (364 mμ) (CCl ₄) ^d	--
Anthracene	3.0 (23 ⁰ C) ^d CH ₂ Cl ₂ , CCl ₄	~550 at 430 mμ (430 mμ) (CCl ₄) ^d	-1.61 (CCl ₄) ^d

1. L.J. Andrews and R.M. Keefer, J. Am. Chem. Soc., **74**, 4500 (1952).
2. J.A.A. Ketelaar, J. Phys. Rad., **15**, 197 (1954).
3. R.M. Keefer and L.J. Andrews, J. Am. Chem. Soc., **77**, 2164 (1955).
4. J. Peters and W.B. Person, J. Am. Chem. Soc., **86**, 10 (1964).